

Potassium

K

General Information

Discovery

Potassium was discovered by Sir Humphry Davy in 1807 in London, by the electrolysis of potassium hydroxide (potash). This was the first metal to be isolated by electrolysis.

Appearance

Potassium is a soft, white metal which is silvery when cut but which rapidly oxidises.

Source

Potassium is the seventh most abundant metal and makes up 2.4% by mass of the earth's crust. Most minerals containing potassium are sparingly soluble and the metal is difficult to obtain from them. Certain minerals, however, such as sylvite, sylvinite and carnallite are found in ancient water beds and potassium salts can be easily recovered from these vast deposits. Potassium hydroxide, potash, is mined in several places including Germany and the USA. Potassium is also found in the ocean in small amounts compared to sodium.

Uses

The greatest demand for potassium is the use of potash in fertilisers. Many other potassium salts are of great importance, including the nitrate, carbonate, chloride, bromide, cyanide and sulphate.

Biological Role

Potassium is essential to life, and non-toxic. One of its natural isotopes is radioactive, and although this radioactivity is mild, it may be one natural cause of genetic mutation in man.

General Information

Potassium is the lightest known metal. It is also one of the most reactive and electropositive of metals, and as it oxidises rapidly in air it must be preserved in a mineral oil such as kerosene. Its reaction with water is vigorous - it catches fire spontaneously and decomposes with the evolution of hydrogen. Potassium and its salts burn with a violet colour.

Physical Information

Atomic Number	19
Relative Atomic Mass ($^{12}\text{C}=12.000$)	39.098
Melting Point/K	336.80
Boiling Point/K	1047
Density/kg m ⁻³	862 (293K)
Ground State Electron Configuration	[Ar]4s ¹
Electron Affinity (M-M ⁻)/kJ mol ⁻¹	43.8

Key Isotopes

Nuclide	³⁹ K	⁴⁰ K	⁴¹ K	⁴² K	⁴³ K
Atomic mass	38.964	39.974	40.962	41.963	42.964
Natural abundance	93.258%	0.0117%	6.730%	0%	0%
Half-life	stable	1.28x10 ⁹ yrs	stable	12 h	22.4 h

Ionisation Energies/kJ mol⁻¹

M - M ⁺	418.8
M ⁺ - M ²⁺	3051.4
M ²⁺ - M ³⁺	4411
M ³⁺ - M ⁴⁺	5877
M ⁴⁺ - M ⁵⁺	7975
M ⁵⁺ - M ⁶⁺	9649
M ⁶⁺ - M ⁷⁺	11343
M ⁷⁺ - M ⁸⁺	14942
M ⁸⁺ - M ⁹⁺	16964
M ⁹⁺ - M ¹⁰⁺	48575

Other Information

Enthalpy of Fusion/kJ mol⁻¹ 2.40

Enthalpy of Vaporisation/kJ mol⁻¹ 79.1

Oxidation States

Main K^I

Others K^{-I}

Covalent Bonds/kJ mol⁻¹

Not applicable