

# Aluminium

# Al

## **General Information**

### **Discovery**

Aluminium was first prepared in an impure form by Hans Christian Oersted in Copenhagen in 1825, and isolated as an element in 1827 by Wohler.

### **Appearance**

Aluminium is a hard and strong, silvery-white metal. An oxide film prevents it from reacting with air and water.

### **Source**

Aluminium is not found free in nature, but is the most abundant metal in the earth's crust (8.1%) in the form of minerals such as bauxite and cryolite. Most commercially produced aluminium is obtained by the Bayer process of refining bauxite. In this process the bauxite is refined to pure aluminium oxide, which is then electrolytically reduced to pure aluminium.

### **Uses**

Aluminium is used in an enormous variety of products, due to its particular properties. It is light, non-toxic, has a high thermal conductivity, has excellent corrosion resistance, and can be easily cast, machined and formed. It is also non-magnetic and non-sparking. It is the second most malleable metal and the sixth most ductile. It is therefore extensively used for kitchen utensils, outside building decoration and in any area where a strong, light, easily constructed material is needed.

The electrical conductivity of aluminium is about 60% that of copper per unit area of cross-section, but it is nevertheless used in electrical transmission lines because of its light weight. Alloys of aluminium with copper, manganese, magnesium and silicon are of vital importance in the construction of aeroplanes and rockets.

Aluminium, when evaporated in a vacuum, forms a highly reflective coating for both light and heat which does not deteriorate as does a silver coating. These aluminium coatings are used for telescope mirrors, in decorative paper, packages, toys and have many other uses.

### **Biological Role**

Aluminium has no known biological role. It can be accumulated in the body from daily intake and has recently been implicated as a potential causative factor in Alzheimer's disease (senile dementia).

## **General Information**

The ancient Greeks and Romans used alum (aluminium oxide) in medicine as an astringent, and in dyeing as a mordant. Sir Humphry Davy proposed the name aluminum for the element, which was undiscovered at the time, and later agreed to change it to aluminium.

Aluminium oxide, alumina, occurs naturally as ruby, sapphire, corundum and emery, and is used in glass-making and refractories.

## Physical Information

Atomic Number	13
Relative Atomic Mass ( $^{12}\text{C}=12.000$ )	26.982
Melting Point/K	933.52
Boiling Point/K	2740
Density/kg m <sup>-3</sup>	2698 (293K)
Ground State Electron Configuration	[Ne]3s <sup>2</sup> 3p <sup>1</sup>
Electron Affinity (M-M <sup>-</sup> )/kJ mol <sup>-1</sup>	44

## Key Isotopes

Nuclide	<sup>26</sup> Al	<sup>27</sup> Al
Atomic mass	25.986	26.982
Natural abundance	0%	100%
Half-life	7.4x10 <sup>5</sup> yrs	stable

## Ionisation Energies/kJ mol<sup>-1</sup>

M - M <sup>+</sup>	577.4
M <sup>+</sup> - M <sup>2+</sup>	1816.6
M <sup>2+</sup> - M <sup>3+</sup>	2744.6
M <sup>3+</sup> - M <sup>4+</sup>	11575
M <sup>4+</sup> - M <sup>5+</sup>	14839
M <sup>5+</sup> - M <sup>6+</sup>	18376
M <sup>6+</sup> - M <sup>7+</sup>	23293
M <sup>7+</sup> - M <sup>8+</sup>	27457
M <sup>8+</sup> - M <sup>9+</sup>	31857
M <sup>9+</sup> - M <sup>10+</sup>	38459

## Other Information

Enthalpy of Fusion/kJ mol<sup>-1</sup> 10.67

Enthalpy of Vaporisation/kJ mol<sup>-1</sup> 290.8

### Oxidation States

Main Al<sup>III</sup>

Others Al<sup>0</sup>, Al<sup>I</sup>

### Covalent Bonds/kJ mol<sup>-1</sup>

Al - H 285

Al - C 225

Al - O 585

Al - F 665

Al - Cl 498

Al - Al 200